

Part A

Charles's law states that the volume (V) of a fixed quantity of gas is directly proportional to its temperature (T) at a constant pressure.

$$V \propto T$$

You can verify this law by plotting the graph of a gas's volume versus its temperature.

To perform this analysis, pump the handle only once so that a fixed number of gas molecules enter the gas chamber. Set the number of "Heavy Species" gas molecules to 100 using the text box given in the tab named **Gas in chamber**. Once the pressure reaches the value of about 0.50 atm, click on the "Pressure" button under the tab **Constant Parameter**, which is at the top right corner of the simulation. Go to the panel named "Tools and Options." Select the ruler by checking off the option in the Measurement Tools. Observe that the height of the cylinder (as measured left to right) does not remain constant because the molecules exert pressure on the walls of the cylinder.

Set the temperature by using the heat control box to add or remove heat as given in the table below.

Temperature (K)	200.	250.	300.	350.
-----------------	------	------	------	------

Use the ruler to find the height of the cylinder as an average between two whole nanometer markings on the ruler. For example, if the value is fluctuating between 6.0 and 7.0 nm, consider 6.5 nm to be the height of the cylinder even if the ruler appears to hover closer to either marking.

Complete the table below with your raw data for the height of the cylinder at each temperature.

Drag the appropriate labels to their respective targets.

3.5

4.5

5.5

6.5

7.5

8.5

Temperature (K)	200.	250.	300.	350.
Height of cylinder (nm)	4.5	5.5	6.5	7.5

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Phet Gas Law Simulation Worksheet Answers

**Paul Baepler, J. D. Walker, D.
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A Guide to Teaching in the Active Learning Classroom Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03 While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment, for administrators planning to create ALCs or experimenting with provisionally designed rooms, and for faculty developers helping teachers transition to using these new spaces.

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