

Name: _____

Date: _____

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pH and pOH Calculations

Find the pH and the pOH of the solution	pH	pOH
1. A 0.023 M solution of perchloric acid (HClO_4)	$\text{pH} = -\log(0.023)$ $\text{pH} = 1.638 \rightarrow 1.6$	$\text{pOH} = 14 - 1.6$ $\text{pOH} = 12.4$
2. A 6.6×10^{-6} M solution of nitric acid (HNO_3)	$\text{pH} = -\log(6.6 \times 10^{-6})$ $\text{pH} = 5.18 \rightarrow 5.2$	$\text{pOH} = 14 - 5.2$ $\text{pOH} = 8.8$
3. A 0.0334 M solution of potassium hydroxide (KOH)	$\text{pH} = 14 - 1.48$ $\text{pH} = 12.52 \rightarrow 12.5$	$\text{pOH} = -\log(0.0334)$ $\text{pOH} = 1.476 \rightarrow 1.48$
4. A 1.0×10^{-5} M solution of hydroarsenic acid (H_3As)	$\text{pH} = -\log(3.0 \times 10^{-5})$ $\text{pH} = 4.52 \rightarrow 4.5$	$\text{pOH} = 14 - 4.5$ $\text{pOH} = 9.5$
5. A 2.23×10^{-12} M solution of aluminum hydroxide $\text{Al}(\text{OH})_3$	$\text{pH} = 14 - 1.17$ $\text{pH} = 12.83 \rightarrow 12.8$	$\text{pOH} = -\log(6.69 \times 10^{-12})$ $\text{pOH} = 1.17$
6. A 0.000901 M solution of magnesium hydroxide $\text{Mg}(\text{OH})_2$	$\text{pH} = 14 - 2.74$ $\text{pH} = 11.26 \rightarrow 11.3$	$\text{pOH} = -\log(0.001802)$ $\text{pOH} = 2.74$
7. A 7.23×10^{-5} M solution of thiosulfic acid ($\text{H}_2\text{S}_2\text{O}_3$)	$\text{pH} = -\log(14.46 \times 10^{-5})$ $\text{pH} = 3.839 \rightarrow 3.84$	$\text{pOH} = 14 - 3.84$ $\text{pOH} = 10.16 \rightarrow 10.2$
8. A 2.58×10^{-5} M solution of oxalic acid ($\text{H}_2\text{C}_2\text{O}_4$)	$\text{pH} = -\log(5.16 \times 10^{-5})$ $\text{pH} = 5.287 \rightarrow 5.29$	$\text{pOH} = 14 - 5.29$ $\text{pOH} = 8.71$
9. A 0.000469 M solution of sodium hydroxide (NaOH)	$\text{pH} = 14 - 3.33$ $\text{pH} = 10.67 \rightarrow 10.7$	$\text{pOH} = -\log(0.000469)$ $\text{pOH} = 3.328 \rightarrow 3.33$
10. A 0.000896 M solution of sulfuric acid (H_2SO_4)	$\text{pH} = -\log(0.01792)$ $\text{pH} = 2.746 \rightarrow 2.75$	$\text{pOH} = 14 - 2.75$ $\text{pOH} = 11.25 \rightarrow 11.3$

Ph And Poh Calculations Answer Key

Marian DeWane, Heather Hattori



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