

13 These calculations have used  $L = 6.02 \times 10^{23}$

(a)  $7.22 \times 10^{22}$  (b)  $3.01 \times 10^{23}$

(c)  $1.20 \times 10^{23}$

14 0.53 mol H

15 0.250 mol

16 (a) 262.87 g mol<sup>-1</sup> (b) 176.14 g mol<sup>-1</sup>

(c) 164.10 g mol<sup>-1</sup> (d) 248.22 g mol<sup>-1</sup>

17 189.1 g

18 1.5 mol

19 0.0074 mol Cl<sup>-</sup>

20  $1.83 \times 10^{23}$  C atoms

21 171 g (integer value because no calculator)

22 10.0 g H<sub>2</sub>O

23 2.0 mol N<sub>2</sub> > 3.0 mol NH<sub>3</sub> > 25.0 mol H<sub>2</sub> > 1.0 mol N<sub>2</sub>H<sub>4</sub>

24 (a) CH (b) CH<sub>2</sub>O

(c) C<sub>10</sub>H<sub>12</sub>O<sub>11</sub> (d) C<sub>2</sub>H<sub>6</sub>

(e) C<sub>2</sub>H<sub>6</sub> (f) CH<sub>3</sub>O

25 Na<sub>2</sub>S<sub>2</sub>O<sub>5</sub>

26 CoSO<sub>4</sub>·7H<sub>2</sub>O

27 C<sub>11</sub>H<sub>15</sub>N

28 NH<sub>3</sub>

29 6.94 L

30 CdS

31 empirical formula CH; molecular formula C<sub>7</sub>H<sub>8</sub>

32 empirical formula H<sub>2</sub>PO<sub>3</sub>; molecular formula H<sub>4</sub>P<sub>2</sub>O<sub>6</sub>

33 C<sub>10</sub>H<sub>12</sub>N<sub>2</sub>P<sub>2</sub>O<sub>11</sub> for both empirical and molecular formulas

34 C<sub>2</sub>H<sub>3</sub>O

35 Let  $y$  = mass of chalk in grams.

moles of chalk used =  $\frac{\text{mass used}}{M_r(\text{CaCO}_3)} \times$

$y \frac{\text{g}}{100.09 \text{ g mol}^{-1}}$

This is the same as the number of moles of carbon atoms used.

Therefore the number of carbon atoms used =

moles of chalk  $\times$  (6.02  $\times$  10<sup>23</sup> mol<sup>-1</sup>) =

$\frac{6.02 \times 10^{23} y}{100.09}$

36 (a) 2.50 mol

(b) 5.63 mol

(c) 665.5 g

37 (a)  $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$

(b) 1.59 g

38 4.355 kg

39 (a)  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$

(b) 92.8%

(c) CaCO<sub>3</sub> is the only source of CO<sub>2</sub>; all the CaCO<sub>3</sub> undergoes complete decomposition; all CO<sub>2</sub> released is captured; heating does not cause any change in the mass of the other minerals present.

40 (a) 65.2 g

(b) 1.3 g H<sub>2</sub>

41 5.23 g C<sub>2</sub>H<sub>5</sub>Cl<sub>2</sub>

42 253.2 g theoretical CaSO<sub>4</sub>; 77.9%

43 3.16 g ester

44 107 g of C<sub>2</sub>H<sub>6</sub> needed

45 (a) 2.40 mol

(b) 0.0110 mol

(c) 44 mol

46 (a) 35.65 dm<sup>3</sup>

(b) 5.7 dm<sup>3</sup>

47 0.652 dm<sup>3</sup>

48 0.138 mol Br<sub>2</sub> and 0.156 mol Cl<sub>2</sub>, so more molecules of Cl<sub>2</sub>

49 0.113 dm<sup>3</sup>

50 0.28 dm<sup>3</sup>

51 90 kPa

52 16 °C

53 3.0 dm<sup>3</sup>

54 2.8 dm<sup>3</sup>

# Pearson Chemistry Chapter 21 2 Workbook Answers

**CO Houle**



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