

7. molar heat of condensation

#### Part B: True-False

8. AT

10. ST

12. NT

#### Part C: Matching

13. a

15. b

17. d

14. c

16. c

#### Part D: Questions and Problems

18. a. endothermic      d. endothermic

    b. endothermic      e. exothermic

    c. exothermic      f. exothermic

19.  $\Delta H = 28.3 \frac{\text{g H}_2\text{O(l)}}{18.0 \text{ g H}_2\text{O(l)}} \times \frac{1 \text{ mol H}_2\text{O(l)}}{1 \text{ mol H}_2\text{O(l)}} \times \frac{6.01 \text{ kJ}}{1 \text{ mol H}_2\text{O(l)}} = 9.45 \text{ kJ}$

20.  $\Delta H = 5.53 \frac{\text{mol NH}_4\text{NO}_3}{25.7 \text{ kJ}} \times \frac{25.7 \text{ kJ}}{1 \text{ mol NH}_4\text{NO}_3} = 142 \text{ kJ}$

#### Section 17.4

#### Part A: Completion

1. sum

2. enthalpy

3. indirectly

4. changed (reversed)

5. standard heat of formation

6. change

7. one

8.  $\Delta H_f^0$

9. zero

10. subtracting

#### Part B: True-False

11. NT

13. AT

15. NT

12. NT

14. NT

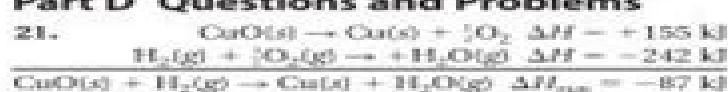
#### Part C: Matching

16. b

18. a

20. c

#### Part D: Questions and Problems



#### 22. $\Delta H_f^0$ (products)

$$= 3 \text{ mol C6H}_6\text{O(l)} \times \frac{-393.5 \text{ kJ}}{1 \text{ mol C6H}_6\text{O(l)}} \\ = -1181 \text{ kJ}$$

#### $\Delta H_f^0$ (reactants)

$$= 3 \text{ mol C6H}_6\text{O(l)} \times \frac{-110.5 \text{ kJ}}{1 \text{ mol C6H}_6\text{O(l)}} + (-822.1 \text{ kJ}) \\ = -1154 \text{ kJ}$$
$$\Delta H^0 = \Delta H_f^0 \text{ (products)} - \Delta H_f^0 \text{ (reactants)}$$
$$= -1181 \text{ kJ} - (-1154 \text{ kJ}) \\ = -27 \text{ kJ}$$

#### Practice Problems

##### Section 17.1

1.  $200.0 \text{ Cal} \times \frac{1000 \text{ cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ cal}} \times \frac{1 \text{ kJ}}{10^3 \text{ J}} \\ = 836.8 \text{ kJ}$

2.  $C = \frac{525.0 \text{ cal}}{25.0 \text{ g} \times 15.0^\circ\text{C}} = 1.40 \text{ cal/(g} \times ^\circ\text{C)}$

3.  $\Delta T = \frac{1255.0 \text{ J}}{100.0 \text{ g} \times 2.1 \text{ J/(g} \times ^\circ\text{C)}} = 5.0^\circ\text{C}$

4.  $q = 100.0 \text{ g} \times 120.0^\circ\text{C} \times 0.90 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \\ = 1.1 \times 10^5 \text{ J}$

##### Section 17.2

1.  $\Delta H = 150.0 \text{ g} \times 4.18 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \times 10^\circ\text{C} \\ = 6.3 \times 10^3 \text{ J} = 6.3 \text{ kJ}$

2.  $\Delta H = 15.0 \frac{\text{g Ca(OH)}_2\text{O(l)}}{74.1 \text{ g Ca(OH)}_2\text{O(l)}} \times \frac{1 \text{ mol Ca(OH)}_2\text{O(l)}}{1 \text{ mol Ca(OH)}_2\text{O(l)}} \\ \times \frac{-63.2 \text{ kJ}}{1 \text{ mol Ca(OH)}_2\text{O(l)}} = -13.2 \text{ kJ}$

3.  $\Delta H = 52.4 \frac{\text{g CH}_3\text{O(l)}}{16.0 \text{ g CH}_3\text{O(l)}} \times \frac{1 \text{ mol CH}_3\text{O(l)}}{16.0 \text{ g CH}_3\text{O(l)}} \\ \times \frac{-290.2 \text{ kJ}}{1 \text{ mol CH}_3\text{O(l)}} = 2.93 \times 10^3 \text{ kJ}$

4.  $4\text{NH}_3 + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO(g)} + 6\text{H}_2\text{O(l)}$   
$$\Delta H = \frac{-226 \text{ kJ}}{1 \text{ mol NH}_3\text{O(l)}} \times 4 \text{ mol NH}_3\text{O(l)} \\ = -904 \text{ kJ}$$

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**DP Hallahan**

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