

13. These calculations have used $L = 6.02 \times 10^{23}$
 (a) 7.22×10^{23} (b) 3.01×10^{23}

14. 0.53 mol H

15. 0.250 mol

16. (a) $262.87 \text{ g mol}^{-1}$ (b) $176.14 \text{ g mol}^{-1}$
 (c) $184.10 \text{ g mol}^{-1}$ (d) $248.22 \text{ g mol}^{-1}$

17. 189.1 g

18. 1.5 mol

19. 0.0074 mol Cl^-

20. 1.83×10^{23} C atoms

21. 17.1 g (integer value because no calculator)

22. 10.0 g H_2O

23. $2.0 \text{ mol N}_2 > 3.0 \text{ mol NH}_3 > 25.0 \text{ mol H}_2 > 1.0 \text{ mol N}_2\text{H}_4$

24. (a) CH (b) CH_2O
 (c) $\text{C}_2\text{H}_4\text{O}_2$ (d) C_2H_6
 (e) C_2H_4 (f) CH_2O

25. $\text{Na}_2\text{S}_2\text{O}_3$

26. $\text{CoSO}_4 \cdot 7\text{H}_2\text{O}$

27. $\text{C}_2\text{H}_5\text{N}$

28. NH_3

29. 6.94 L

30. CoS

31. empirical formula CH ; molecular formula C_2H_4

32. empirical formula H_2PO_4^- ; molecular formula $\text{H}_4\text{P}_2\text{O}_7$

33. $\text{C}_{12}\text{H}_{16}\text{N}_4\text{P}_2\text{O}_{10}$, for both empirical and molecular formulas

34. $\text{C}_2\text{H}_4\text{O}$

35. Let $y = \text{mass of chalk in grams}.$
 $\text{moles of chalk used} = \frac{\text{mass used}}{M(\text{CaCO}_3)} \times$
 $y \frac{\text{g}}{100.09 \text{ g mol}^{-1}}$

36. This is the same as the number of moles of carbon atoms used.
 Therefore the number of carbon atoms used =
 $\text{moles of chalk} \times (6.02 \times 10^{23} \text{ mol}^{-1}) =$
 $6.02 \times 10^{23} y$

37. (a) $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$
 (b) 1.59 g

38. 4.355 kg

39. (a) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
 (b) 92.8%
 (c) CaCO_3 is the only source of CO_2 ; all the CaCO_3 undergoes complete decomposition; all CO_2 released is captured; heating does not cause any change in the mass of the other minerals present.

40. (a) 65.2 g (b) 1.3 g H_2

41. 5.23 g $\text{C}_2\text{H}_4\text{O}_2$

42. 253.2 g theoretical CaSO_4 ; 77.9%

43. 3.16 g ester

44. 107 g of C_2H_6 needed

45. (a) 2.40 mol (b) 0.0110 mol

46. (a) 35.85 dm^3 (b) 5.7 dm^3

47. 0.652 dm^3

48. 0.138 mol Br_2 and 0.156 mol Cl_2 , so more molecules of Cl_2

49. 0.113 dm^3

50. 0.28 dm^3

51. 90 kPa

52. 16 °C

53. 3.0 dm^3

54. 2.8 dm^3

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Answers

Alan Hess

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