

Acids and bases

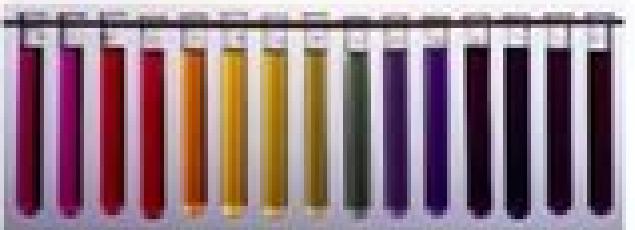
Acids are substances that:

- Have a pH below 7.
- Produce H^+ ions in aqueous solutions.

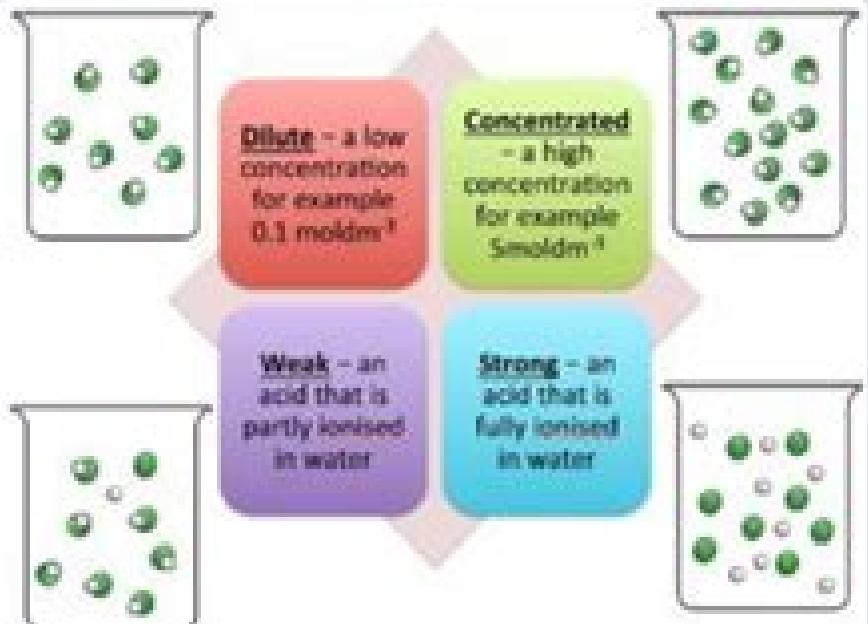


Bases are substances that:

- React with the H^+ ions in acids.
- Are called **alkalis** if they dissolve in water.
- Is a proton acceptor (H^+ ions).
- Produce OH^- ions in aqueous solutions.



3



The pH scale and universal indicator

- The pH scale (0 to 14) measures the acidity or alkalinity of a solution.
- pH can be measured using universal indicator or a pH probe.



4

Experimental techniques to determine acid or base strengths

- Conductivity** - strong acids/bases will be better conductors. This is due to more ions being present in the solution. E.g. If we tested the conductivity of 1 mol dm^{-3} ethanoic acid compared to 1 mol dm^{-3} of hydrochloric acid, the hydrochloric acid would be more conductive. This can be measured with a pH meter or electrodes/conductivity meter.
- pH** - Strong acids are pH 0-3, weak acids pH 4-6, weak bases pH 8-10, strong bases pH 11-14. This can be measured with universal indicator solution/paper or a pH meter.
- Rates of reaction** - The reaction of an acid depends on $[\text{H}^+]$, so a greater rate means a stronger acid (e.g. reactions with metals/carbonates).

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A. Thorpe, J. Hammond, A. Baldwin, D. Thomson, D. Root, M. Murray

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