

Name Key Class \_\_\_\_\_ Date \_\_\_\_\_

## Nuclear Decay

### Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, positron, and gamma radiation.

Name	Alpha	Beta	Positron	Gamma
Greek letter	$\alpha$	$\beta^-$	$\beta^+$	$\gamma$
Symbol	${}^4_2\text{He}$	${}^0_{-1}e$	${}^0_{+1}e$	N/A
Composition	$2p^+$ , $2n^0$	electron	anti-electron	EM wave
Charge	$2+$	$1-$	$1+$	$0$
Stopped by	paper, skin, clothing	wood, glass	wood, glass	thick concrete or lead

### Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.



A. alpha emission



B. beta emission



C. electron capture



D. positron emission

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha ( $\alpha$ ) emission.



7. Decay of iodine-126 by positron ( $\beta^+$ ) emission.



8. How are the atomic number and mass number of a nuclide affected by positron emission?

**Atomic number decreases by one. Mass number stays the same.**

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

**Both atomic number and mass number stay the same since gamma radiation has no mass or charge.**

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

**The reactants have a slightly greater mass. In a nuclear reaction, a small amount of mass is converted to energy according to the equation  $E = mc^2$ . The difference in mass is referred to as the mass defect.**

# **Nuclear Decay Chemistry If8766 Answers**

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Quality of Life**



## **Nuclear Decay Chemistry If8766 Answers:**

*Principles of Nuclear Chemistry* Peter A. C. McPherson, 2017      **Nuclear Decay Scheme Analysis and Characterization Studies of (d, Alpha) Reaction Products** Donald Glenn Gardner, 1957      *Nuclear Decay Scheme Analysis and Characterization Studies of (d, Alpha) Reaction Products* Donald Glenn Gardner, University of Michigan. Department of Chemistry, 1957      **Nuclear Decay Data** International Committee for Radionuclide Metrology. Symposium on Nuclear Decay Data, 1990      *The Heart of the Matter* Victor E. Viola, 1973      Nuclear Chemistry Maheshwar Sharon, Madhuri Sharon, 2009-09-25 Concentrating on techniques for the detection and measurement of radioactivity this book offers a guide to selecting the type of counter type of source sample duration for which the counting must be made and the radiation emitted by the isotope for its efficient detection It introduces a novel concept to explain not only the decay processes but also the selection of counting procedures for detecting and measuring radioactivity The author builds up the foundation from the nature of the interaction of radiation with matter He also highlights the differences between an ordinary chemical laboratory and a radiochemical one      **Nuclear and Radiochemistry** G. Friedlander, 1966      **SPALLDKZ** Roger Andrew Gard, 1999      **A Spectroscopic Approach to Nuclear Decay by Electron Capture** Marc Merstorf, 2022\*      **Nuclear Decay Scheme Analysis and Characterization Studies of (d, Alpha) Reaction Products by Donald Glenn Gardner** Donald Glenn Gardner, 1957      **Nuclear Decay Scheme Studies** Jackie Devon Woodward, 1968      *Determination of Nuclear Decay Schemes by the Coincidence Method* James Arthur Cooley, 1952      **Nuclear Decay D.** Russell Humphreys, Institute for Creation Research, 2002      **Coincidence Measurements in Nuclear Decay Scheme Studies** John Peter Unik, 1960      **Nuclear Decay Data**, 1990      **Handbook of Nuclear Decay Modes** Dorin N. Poenaru, 1993-11-01      *Theory of Alpha Decay*, 1962 The relationship between nuclear reactions and nuclear decay derived by F T Smith is presented and illustrated by a simple example auth      Evidence for Variation in Nuclear Decay Rates, 2012      **Recommended Nuclear Decay Data** Andy Pearce, National Physical Laboratory (Great Britain). Division of Quality of Life, 2008      **Nuclear Decay Curve Analysis** Ditmar Bock, 1969

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kinetic energy of a 1-kilogram ball is thrown into the air with an initial velocity of 30 m/sec?  $KE = \frac{1}{2} m v^2$   $\frac{1}{2} (1 \text{ kg}) \dots$  Kinetic Energy (KE) =  $\frac{1}{2}$  mass times velocity squared Potential and Kinetic Energy Worksheet. Kinetic Energy (KE) =  $\frac{1}{2}$  mass times velocity squared.  $KE = \frac{1}{2} m v^2$ . Potential Energy (PE) = mass times the acceleration ... Kinetic and potential energy worksheet answer key o myaiu kinetic and potential energy worksheet classify the following as type of potential energy or kinetic energy (use the letters or bicyclist pedaling up ... Kinetic and Potential Energy Worksheet Walkthrough - YouTube kinetic and potential energy worksheet Flashcards A. How much kinetic energy does the ball have? B. How much potential energy does the ball have when it reaches the top of the ascent? KINETIC AND POTENTIAL ENERGY WORKSHEET Answer the following:

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