

Nuclear Reactions

Nuclear reactions are characterized by the following:

- They involve changes to the nucleus
 - Number of protons or neutrons in the nucleus change
- Energy changes are enormous
- Not concerned with valence electrons.
- Changes in temperature and pressure have no effect on the rate of the reaction
- Small particles are important.

Important subatomic particles

Prototypes: \mathcal{P} \mathcal{H}

neutrons: $\frac{1}{2}$ me

Alpha Particle: ${}^4_2\text{He}$ 2 protons
2 neutrons

beta particle: ${}_{-1}^0 \text{B} \rightarrow {}_{-1}^0 \text{e}$ - high energy electrons emitted from the nucleus

position: ${}^0_+ \beta^- = {}^0_+ e^-$ - positive electron: what gives a proton its positive charge

Gamma Ray : γ - high energy photons

Definitions

Isotopes – same number of protons but different mass number (mass number = # protons + # neutrons)
 ^{12}C , ^{13}C , ^{14}C are each an isotope of C (How many neutrons does each have?)

Nuclide (pronounced NOO-uh-ide) — a nucleus with a specific number of protons and neutrons.

Radioactive - Changes in matter resulting from SPONTANEOUS changes in the composition of the nucleus.

- Such changes are often accompanied by the creation of the subatomic particles above.
- Changes to the nucleus often result in the conversion of one element into another.

Radioisotope – nuclides that are radioactive.



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