

Name: KEY

15 Pts.

## Chemistry Practice: Writing Chemical Formulas

Write a chemical formula for each substance.

- |  |                                |  |                           |
|--|--------------------------------|--|---------------------------|
| 1. <u>NaCl</u>   | sodium chloride                | 39. <u>Ca(NO<sub>3</sub>)<sub>2</sub></u>                          | calcium nitrate           |
| 2. <u>N<sub>2</sub>O<sub>5</sub></u>                               | dinitrogen pentoxide           | 40. <u>SO<sub>3</sub></u>  | sulfur trioxide           |
| 3. <u>H<sub>2</sub>S</u>   | hydrosulfuric acid             | 41. <u>KCN</u>   | potassium cyanide         |
| 4. <u>K<sub>2</sub>SO<sub>4</sub></u>                              | potassium sulfate              | 42. <u>Pb(NO<sub>3</sub>)<sub>2</sub></u>                          | lead(II) nitrate          |
| 5. <u>H<sub>2</sub>C<sub>2</sub>O<sub>4</sub></u>                  | oxalic acid                    | 43. <u>H<sub>2</sub>S</u>  | hydrogen sulfide          |
| 6. <u>Ag<sub>2</sub>C<sub>2</sub>H<sub>3</sub>O<sub>2</sub></u>    | silver acetate                 | 44. <u>CoCl<sub>3</sub></u>  | cobalt(III) chloride      |
| 7. <u>Cr(NO<sub>3</sub>)<sub>3</sub></u>                           | chromium(III) nitrate          | 45. <u>SF<sub>6</sub></u>  | sulfur hexafluoride       |
| 8. <u>HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub></u>                 | acetic acid                    | 46. <u>Ca<sub>3</sub>N<sub>2</sub></u>                             | calcium nitride           |
| 9. <u>(NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub></u>               | ammonium carbonate             | 47. <u>CuI</u>   | copper(I) iodide          |
| 10. <u>Ca(OH)<sub>2</sub></u>                                      | calcium hydroxide              | 48. <u>SiO<sub>2</sub></u>   | silicon dioxide           |
| 11. <u>H<sub>2</sub>C<sub>4</sub>H<sub>4</sub>O<sub>6</sub></u>    | tartaric acid                  | 49. <u>Sn(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>4</sub></u> | tin(IV) acetate           |
| 12. <u>Hg(NO<sub>3</sub>)<sub>2</sub></u>                          | mercury(II) nitrate            | 50. <u>CCl<sub>4</sub></u>   | carbon tetrachloride      |
| 13. <u>N<sub>2</sub>O</u>  | dinitrogen monoxide            | 51. <u>CuS</u>   | copper(II) sulfide        |
| 14. <u>Fe<sub>2</sub>O<sub>3</sub></u>                             | iron(III) oxide                | 52. <u>Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub></u>              | lead(II) phosphate        |
| 15. <u>Pb(ClO<sub>3</sub>)<sub>2</sub></u>                         | lead(II) chlorate              | 53. <u>XeCl<sub>4</sub></u>  | xenon tetrachloride       |
| 16. <u>(NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub></u>              | ammonium phosphate             | 54. <u>Rb<sub>2</sub>O</u>   | rubidium oxide            |
| 17. <u>ZnCl<sub>2</sub></u>  | zinc chloride                  | 55. <u>MgSe</u>  | magnesium selenide        |
| 18. <u>Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub></u>              | calcium phosphate              | 56. <u>NH<sub>4</sub>Cl</u>  | ammonium chloride         |
| 19. <u>OF<sub>2</sub></u>  | oxygen difluoride              | 57. <u>Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub></u> | iron(III) acetate         |
| 20. <u>NaN<sub>3</sub></u>   | sodium azide                   | 58. <u>K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub></u>                | potassium dichromate      |
| 21. <u>Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></u>              | iron(III) sulfate              | 59. <u>PBr<sub>3</sub></u>   | phosphorous tribromide    |
| 22. <u>H<sub>3</sub>AsO<sub>3</sub></u>                            | arsenous acid                  | 60. <u>Na<sub>3</sub>PO<sub>3</sub></u>                            | sodium phosphite          |
| 23. <u>Cr<sub>2</sub>O<sub>3</sub></u>                             | chromium(III) oxide            | 61. <u>Na<sub>3</sub>PO<sub>4</sub></u>                            | sodium phosphate          |
| 24. <u>N<sub>2</sub>O<sub>4</sub></u>                              | dinitrogen tetroxide           | 62. <u>Hg(NO<sub>3</sub>)<sub>2</sub></u>                          | mercury(II) nitrate       |
| 25. <u>NH<sub>4</sub>NO<sub>3</sub></u>                            | ammonium nitrate               | 63. <u>LiHCO<sub>3</sub></u>                                       | lithium bicarbonate       |
| 26. <u>AsBr<sub>3</sub></u>  | gold(III) bromide              | 64. <u>CrF<sub>3</sub></u>   | chromium(III) fluoride    |
| 27. <u>CO</u>  | carbon monoxide                | 65. <u>PbI<sub>2</sub></u>   | lead(II) iodide           |
| 28. <u>K<sub>2</sub>CO<sub>3</sub></u>                             | potassium carbonate            | 66. <u>H<sub>2</sub>SO<sub>3</sub></u>                             | sulfurous acid            |
| 29. <u>HIO<sub>3</sub></u>   | iodic acid                     | 67. <u>SnF<sub>2</sub></u>   | tin(II) fluoride          |
| 30. <u>CsCl</u>  | cesium chloride                | 68. <u>HgCrO<sub>4</sub></u>                                       | mercury(II) chromate      |
| 31. <u>Ni(MnO<sub>4</sub>)<sub>2</sub></u>                         | nickel(II) permanganate        | 69. <u>KNO<sub>3</sub></u>   | potassium nitrate         |
| 32. <u>Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></u>              | aluminum sulfate               | 70. <u>SrCl<sub>2</sub></u>  | strontium chloride        |
| 33. <u>Al<sub>2</sub>(SO<sub>3</sub>)<sub>3</sub></u>              | aluminum sulfite               | 71. <u>P<sub>4</sub>O<sub>10</sub></u>                             | tetraphosphorous decoxide |
| 34. <u>Ba(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></u> | barium acetate                 | 72. <u>KNO<sub>3</sub></u>   | potassium nitrate         |
| 35. <u>Mn(OH)<sub>3</sub></u>                                      | manganese(III) hydroxide       | 73. <u>KNO<sub>2</sub></u>   | potassium nitrite         |
| 36. <u>KH<sub>2</sub>PO<sub>4</sub></u>                            | potassium dihydrogen phosphate | 74. <u>K<sub>3</sub>N</u>  | potassium nitride         |
| 37. <u>HF</u>  | hydrofluoric acid              | 75. <u>CaO</u>   | calcium oxide             |
| 38. <u>Al(BrO<sub>3</sub>)<sub>3</sub></u>                         | aluminum bromate               | 76. <u>Fe(IO<sub>4</sub>)<sub>2</sub></u>                          | iron(II) periodate        |

# Nomenclature And Formula Lab Answer Key

**J. A. Beran, Mark Lassiter**



## Nomenclature And Formula Lab Answer Key:

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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

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**Laboratory Manual for Fundamentals of Chemistry 3/E** Jo A. Beran, 1988

**Laboratory Experiments for General, Organic & Biochemistry** Frederick A. Bettelheim, Joseph Marvin Landesberg, 1997

The Stirring Rod, 1920

*Laboratory Manual for Fundamentals of General, Organic, and Biological Chemistry* John R. Holum, Ruth C. Denison, 1978

*General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

*Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science*, 2003-11

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## **Table of Contents Nomenclature And Formula Lab Answer Key**

1. Understanding the eBook Nomenclature And Formula Lab Answer Key
  - The Rise of Digital Reading Nomenclature And Formula Lab Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Nomenclature And Formula Lab Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nomenclature And Formula Lab Answer Key
  - User-Friendly Interface

4. Exploring eBook Recommendations from Nomenclature And Formula Lab Answer Key
  - Personalized Recommendations
  - Nomenclature And Formula Lab Answer Key User Reviews and Ratings
  - Nomenclature And Formula Lab Answer Key and Bestseller Lists
5. Accessing Nomenclature And Formula Lab Answer Key Free and Paid eBooks
  - Nomenclature And Formula Lab Answer Key Public Domain eBooks
  - Nomenclature And Formula Lab Answer Key eBook Subscription Services
  - Nomenclature And Formula Lab Answer Key Budget-Friendly Options
6. Navigating Nomenclature And Formula Lab Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Nomenclature And Formula Lab Answer Key Compatibility with Devices
  - Nomenclature And Formula Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Nomenclature And Formula Lab Answer Key
  - Highlighting and Note-Taking Nomenclature And Formula Lab Answer Key
  - Interactive Elements Nomenclature And Formula Lab Answer Key
8. Staying Engaged with Nomenclature And Formula Lab Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Nomenclature And Formula Lab Answer Key
9. Balancing eBooks and Physical Books Nomenclature And Formula Lab Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Nomenclature And Formula Lab Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Nomenclature And Formula Lab Answer Key
  - Setting Reading Goals Nomenclature And Formula Lab Answer Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Nomenclature And Formula Lab Answer Key
  - Fact-Checking eBook Content of Nomenclature And Formula Lab Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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