

1. What is a wave? Label all of its parts! Wavelength, frequency, amplitude, crest, trough
2. What is the symbol for wavelength?
3. What are the units for wavelength?
4. What is the definition for wavelength?
5. What is frequency?
6. What are the units for frequency?
7. What is the symbol for frequency?
8. What is the electromagnetic spectrum?
9. What is electromagnetic radiation?
10. What are the seven major types of electromagnetic radiation?
11. What is the relationship between frequency and wavelength?
12. What is the relationship between frequency and energy?
13. What is the equation used to solve frequency and wavelength problems?
14. What is the speed of light?
15. Explain the photoelectric effect? (you can use the video we watched to explain it)
16. Who suggested the theory of quanta?
17. What is a quantum?
18. What is Planck's constant?
19. What is the equation for energy?
20. Who changed the equation for energy and how did the change it?
21. What is a photon?
22. What is the continuous spectrum?
23. What is the emission line spectrum?
24. How does an electron move from the ground state to the excited state?
25. What is the difference between absorption and emission?
26. When does an atom gain energy, absorption or emission?
27. Explain Bohr's model of the atom.
28. What is the frequency in of light that has a wavelength of 310nm?
29. What is the wavelength of electromagnetic radiation if its frequency is 3.2×10^8 MHz?
30. What is the frequency in Megahertz of orange light which has a wavelength of 3.25×10^5 cm?
31. What is the energy of a gamma ray that has a frequency of 4.66×10^{19} hertz?
32. What is the frequency in hertz of a microwave if it produces 1.36×10^{14} J?

Quantum Theory Study Guide

Jerry B. Marion



Quantum Theory Study Guide:

Introducing Quantum Theory J.P. McEvoy, Oscar Zarate, 2014-06-05 Quantum theory confronts us with bizarre paradoxes which contradict the logic of classical physics At the subatomic level one particle seems to know what the others are doing and according to Heisenberg's uncertainty principle there is a limit on how accurately nature can be observed And yet the theory is amazingly accurate and widely applied explaining all of chemistry and most of physics Introducing Quantum Theory takes us on a step by step tour with the key figures including Planck Einstein Bohr Heisenberg and Schrodinger Each contributed at least one crucial concept to the theory The puzzle of the wave particle duality is here along with descriptions of the two questions raised against Bohr's Copenhagen Interpretation the famous dead and alive cat and the EPR paradox Both remain unresolved

Introducing Quantum Theory Joseph P. McEvoy, Oscar Zarate, 2004 Outlines quantum theory and the key players including Planck Einstein Bohr Heisenberg and Shrodinger

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

SCM Studyguide to Science and Religion Jean Dorricott, 2005 This undergraduate level one textbook provides an introduction to the apparently incompatible subjects of religion and science Each chapter contains references for finding out more about particular arguments be they scientific or religious areas for discussion Where particularly difficult concepts are referred to in the body of the text further explanations are provided in boxed sections

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