

# Science

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(Chapter - 6) (Periodic Classification of Elements) (NIOS)

(Class - X)

## Intext Questions 6.1

### Question 1:

Elements A, B and C constitute a Dobereiner's triad. The atomic mass of A is 20 and that of C is 40. Predict the atomic mass of B.

#### Answer 1:

$$\text{Atomic mass of B} = \frac{(\text{Atomic mass of A} + \text{Atomic mass of C})}{2} = \frac{20 + 40}{2} = \frac{60}{2} = 30$$

Hence, atomic mass of B is 30.

### Question 2:

Which property of atoms was used by Mendeleev to classify the elements?

#### Answer 2:

Mendeleev used the atomic mass of an atom in the classification of elements.

### Question 3:

In Mendeleev's periodic classification, whether chemically similar elements are placed in a group or in a period?

#### Answer 3:

In the Mendeleev periodic table, elements that are chemically similar are placed in groups.

### Question 4:

Mendeleev's periodic table had some blank spaces. What did they signify?

#### Answer 4:

Some of the blank spaces in Mendeleev's periodic table are for elements that have yet to be discovered.

### Question 5:

Explain any three defects of Mendeleev's periodic table.

#### Answer 5:

- Isotopes should be given separate place because they have different atomic mass and character.
- The increasing order of atomic weight is not maintained.
- Some elements in the same group differ in their properties.
- The position of hydrogen is not justified.

## Intext Questions 6.2

### Question 1:

Give any two defects of Mendeleev's periodic table which has been removed in modern periodic table. How were they removed?

#### Answer 1:

Odd Pair: When the elements are arranged in the order of their increasing atomic numbers, the three oddities automatically disappear because the atomic number of the first element is less than that of the second element although it is in the reverse order of their atomic masses.

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**VM Jensen**

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